ABBREVIATIONS AND NOTATIONS

activ.	activated
amorph.	amorphous
anhydr.	anhydrous
bl.	blue
blk.	black
bp.	boiling point at 760 mm pressure
brn.	brown
ca.	*circa* (approximately)
cat.	catalyst
cl.	colorless
conc.	concentrated
cryst.	crystal(line)
d	density of liquid *or* solid substance relative to water at 4°C
dild.	diluted
dk-	dark
e^-	electron
edta	ethylenediaminetetraacetate ion, $C_{10}H_{12}O_8N_2^{4-}$
e.g.	*exempli gratia* (for example)
el.	electric(al)
en	ethylenediamine, $C_2H_8N_2$
ether	diethyl ether
(g)	gaseous (*subscript*)
gas.	gaseous
g/l	gram per litre
grn.	green
Hdmg	dimethylglyoximate ion, $C_4H_7O_2N_2^-$
higher	greater than *or* equal to *t* (specified temperature value)
hydr.	(cristal) hydrate, hydrated
immisc.	immiscible
impur.	impurity, impurities
insol.	insoluble
k	mass solubility coefficient in g/100 g H_2O
K_a	acidity constant at 25°C, $pK_a = -\log K_a$

K_b	basicity constant at 25°C, $pK_b = -\log K_b$
K_c	concentration equilibrium constant, $pK_c = -\log K_c$
K_{inst}	instability constant of complex in aqueous solution at 25°C, $pK_{inst} = -\log K_{inst}$
K_{solv}	ion product constant of solvent
K_{st}	stability constant of complex in aqueous solution at 25°C
K_w	ion product constant of water
L	solubility product constant
(l)	liquid (*subscript*)
lower	less than *or* equal to *t* (specified temperature value)
lq.	liquid
lt-	light
M_r	relative molecular mass (formula weight)
misc.	miscible
mp.	melting point
n/react.	does not react
ntp.	normal temperature and pressure
p	under excess pressure
pH	$-\log[H_3O^+]$ value (in aqueous solution)
py	pyridine, C_5H_5N
react.	reacts (*viz.*, completely hydrolyzes, oxidizes, *or* reduces by water)
room *t*	room temperature, in room conditions
(s)	solid (*subscript*)
satd.	saturated
sk-bl.	sky-blue
sld.	solid
sl. sol.	slightly soluble
sol.	readily soluble
soln.	aqueous solution
t	temperature
t_{dec}	decomposition temperature
t_{dehydr}	dehydration temperature
t_{depol}	depolymerization temperature
t_{polym}	polymerization temperature

t_{soft}	softening temperature	α	equilibrium degree of forward reacting
t_{subl}	sublimation temperature	ρ	density of gaseous substance in g/l
UV irradn.	ultraviolet irradiation	τ	sluggish reacting
ν	volume solubility coefficient in ml(ntp)/100 g H_2O	$\varphi°$	standard electrode potential in aqueous solution (in volt, V)
vac.	vacuum	$>$	greater than
v. dild.	very diluted	$\geq$	greater than or equal to
vitr.	vitreous	$<$	less than
viz.	videlicet (that is)	$\leq$	less than or equal to
wh.	white	$\ll$	much less than
yel.	yellow		